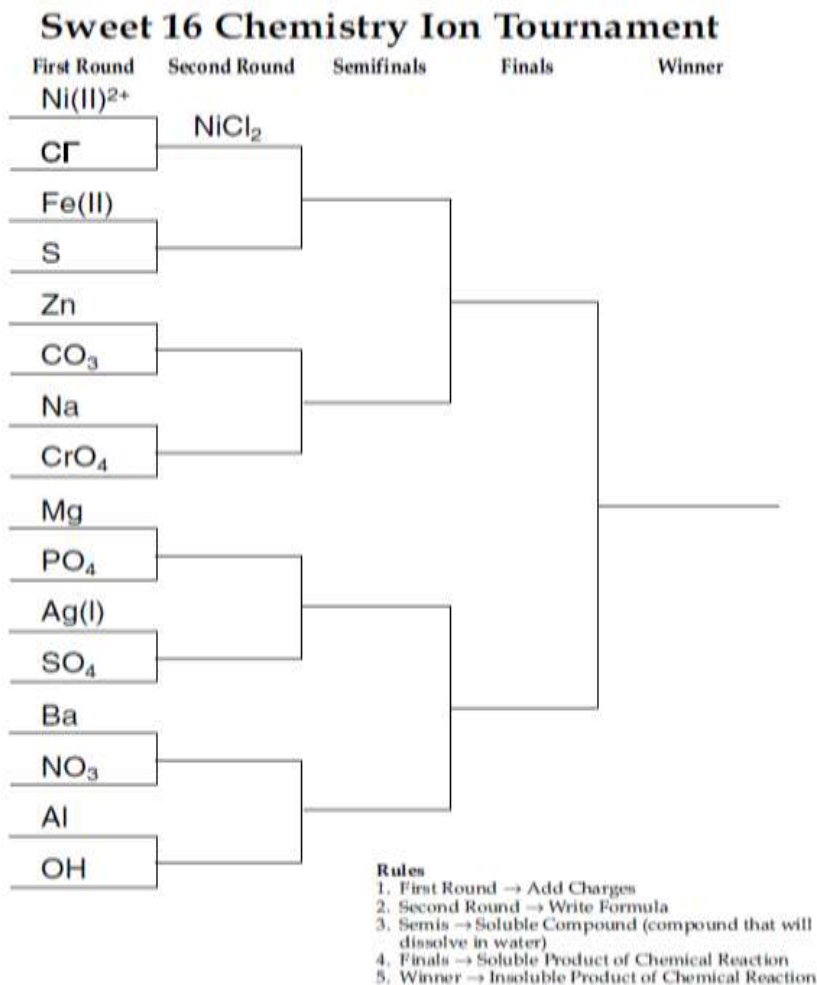


Pre-AP Test 3 S2 Review

Precipitation Rxn, Net ionic Equations, Solutions, Concentrations

You may complete of this paper, but use additional paper when needed. Please use your notes when completing

1. A solution is made of two parts, what are they?
2. There three types of solutions we learned about, two are stable, one id not.
 - a. What are the two sable solution types and how do they differ?
 - b. What is the third type of solution?
3. How are molarity and molality different?
4. Are both molarity and molality temperature dependent, why or why not?
5. What is a precipitate?
6. To form a net ionic equation, what is canceled out of the total ionic equations?



Practical:

Complete the calculations and answer each question, #1-9, then go to the lab and complete.

1. You need to prepare a 100mL 0.1M solution of calcium chloride, how would you do this? Show all calculations and explain steps.

2. You need to prepare a 100mL 0.1M solution of sodium carbonate, how would you do this? Show all calculations and explain steps

3. What is the balanced equation when reacting sodium carbonate with calcium chloride?

4. Is a precipitate formed?

5. Write the balanced precipitation rxn between sodium carbonate and calcium chloride.

6. Write the total ionic equation.

7. Write the net ionic equation.

8. How many moles of soluble product are produced? (Assume stoichiometric equivalent reactant concentrations, no limiting reagent)?

9. What is the molarity of the soluble product (assume all volumes are summative)?

10. Now with a partner, make both solutions and mix them in a separate container. Thoroughly clean all glassware between uses, dispose of in waste hood labeled beaker.

11. How could you recover the soluble product and precipitate separately from your solution?